Acids & Bases

Reading: Chapter 6—Neutralizing the Threat of Acid Rain

What are Acids & Bases

Neutralization

pH & Acid Rain

Sulfur Dioxide and the Combustion of Coal

Nitrogen Oxides and the Acidification of Los Angeles

Acid Deposition & Impact on Materials

Control Strategies

The Politics of Acid Rain



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Acids: Chemical Definition

In general:
$$HX + H_2O \rightarrow H_3O^+ + X^-$$

(we won't use H⁺ for acid although the book does)

Acids produce **hydronium cation** (H₃O⁺) when dissolved in H₂O solution

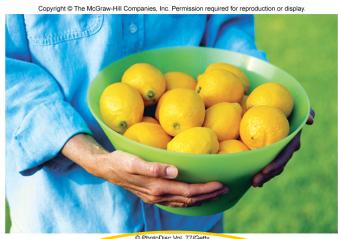
Noble gas configuration maintained

$$CO_2 + H_2O \longrightarrow H_2CO_3$$
 $H_0 \longrightarrow OH$ Carbonic acid

 $H_2CO_3 + H_2O \longrightarrow H_3O^+ + HCO_3^-$ bicarbonate ion

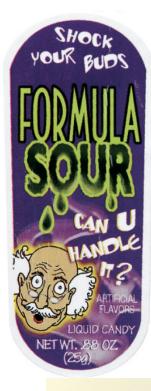
 $HCO_3^- + H_2O \longrightarrow H_3O^+ + CO_3^{2-}$ carbonate polyatomic ions that we talked about in Chapter 5

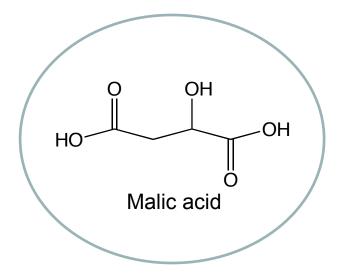
Some Organic Acids



Ascorbic acid (vitamin C): antioxidant

Acids are sour...





Ingred.: cane sugar, water, malic acid, citric acid, natural and artificial flavors, sodium benzoate, FD&C Yellow 5, Yellow 6, Blue 1, Yellow 5 Lake, Blue 2 Lake, Red 40 Lake Mfg. by Squire Boone Village New Albany, IN 47150 For nutrition info, call 1-800-234-1804

Bases: Chemical Definition

Bases are anti-acids (antacids)

Bases produce **hydroxide anion** (OH⁻) when dissolved in H₂O solution

In general: YOH \rightarrow Y⁺(ag) + OH⁻(ag)

NaOH \rightarrow Na⁺(ag) + OH⁻(ag)

 $NH_3(aq) + H_2O \rightarrow NH_4^+(aq) + OH^-(aq)$

Bases are bitter



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Equilibria

$$2H_2O \rightleftharpoons H_3O^+ + OH^-$$

- reaction proceeds in both directions at the same time
- adding more of a reagent will cause it to react faster
- this causes the equilbrium to shift towards the other side

$$K_{eq} = \frac{[Products]}{[Reactants]}$$

$$K_{\mathsf{H}_2\mathsf{O}} = \frac{\left[\mathsf{H}_3\mathsf{O}^+\right]\left[\mathsf{O}\mathsf{H}^-\right]}{\left[\mathsf{H}_2\mathsf{O}\right]} \Rightarrow K_{\mathsf{H}_2\mathsf{O}}\left[\mathsf{H}_2\mathsf{O}\right] = \left[\mathsf{H}_3\mathsf{O}^+\right]\left[\mathsf{O}\mathsf{H}^-\right] \Rightarrow 1x10^{-14} = \left[\mathsf{H}_3\mathsf{O}^+\right]\left[\mathsf{O}\mathsf{H}^-\right]$$

$$[H_2O] = \frac{\text{moles } H_2O}{\text{L solution}}$$

$$1 L H_2O \times \frac{1000 \text{ g H}_2O}{1 L H_2O} \times \frac{\text{mole H}_2O}{18 \text{ g H}_2O} = 55.6 \text{ mole H}_2O \Rightarrow [H_2O] \text{ is 55.6 M}$$

What is pH and Why Does it Matter?

pH stands for power of hydrogen pH = $-log[H_3O^+]$ (negative of the power of 10) $[H_3O^+] = 10^{-pH}$

if $[H_3O^+]=1x10^{-4}$ then pH= 4 if $[H_3O^+]=1x10^{-9}$ then pH= 9 if $[H_3O^+]=1x10^{-7}$ then pH= 7 Congratulations! You are
"Perfectly Norm.

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The perfect start for creating whatever look you want today.

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- pH-balanced formula enriched with Chamomile helps maintain your hair's natural moisture balance

Acids & bases can break bonds in molecules of life

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