



Acids & Bases

Reading: Chapter 6—Neutralizing the Threat of Acid Rain

What are Acids & Bases

Neutralization

pH & Acid Rain

Sulfur Dioxide and the
Combustion of Coal

Nitrogen Oxides and the
Acidification of Los Angeles

Acid Deposition & Impact on
Materials

Control Strategies

The Politics of Acid Rain

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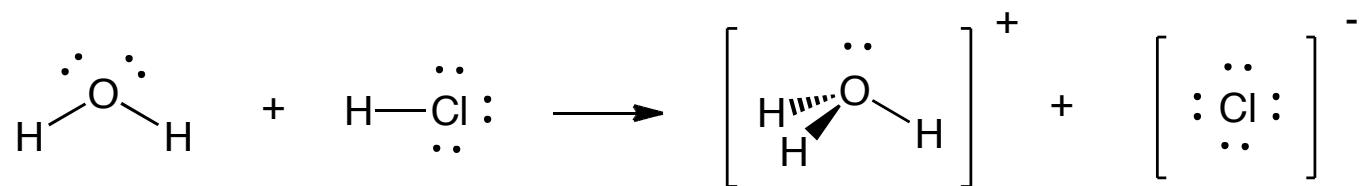
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Acids: Chemical Definition

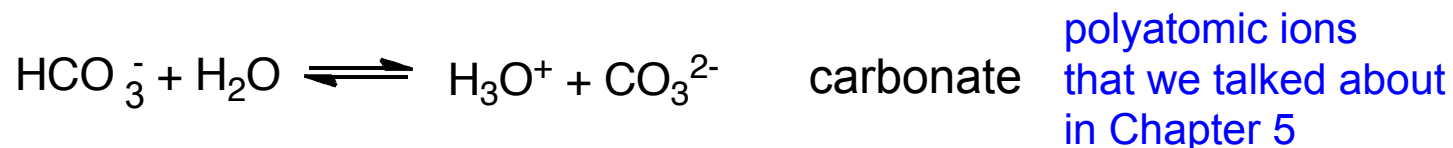
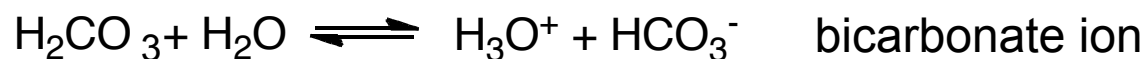
In general: $\text{HX} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{X}^-$

(we won't use H^+ for acid although the book does)

Acids produce **hydronium cation** (H_3O^+) when dissolved in H_2O solution



Noble gas configuration maintained

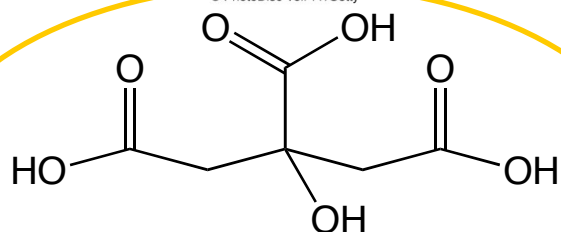


Some Organic Acids

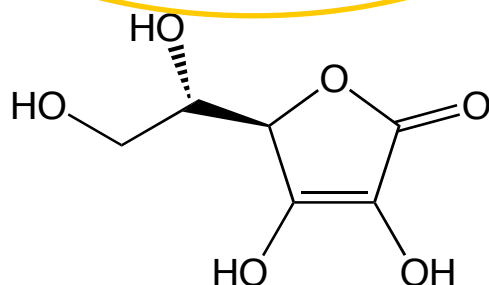
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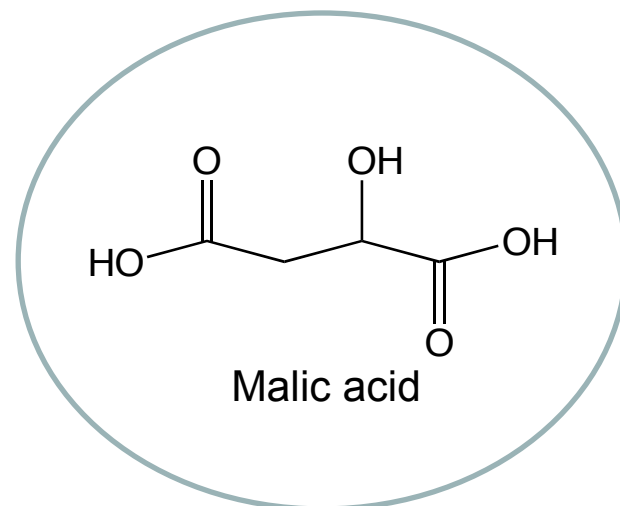
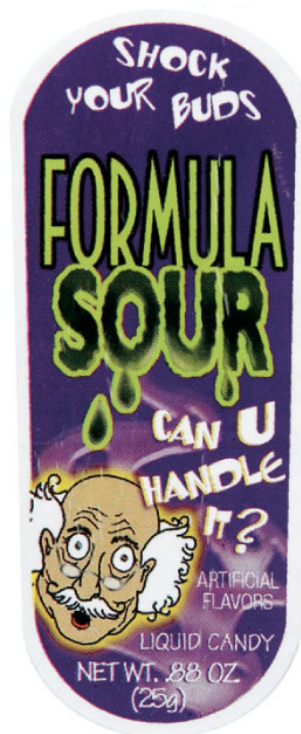


Citric acid: (prevents scurvy)



Ascorbic acid (vitamin C): antioxidant

Acids are sour...



Malic acid

Ingred.: cane sugar, water, malic acid, citric acid, natural and artificial flavors, sodium benzoate, FD&C Yellow 5, Yellow 6, Blue 1, Yellow 5 Lake, Blue 2 Lake, Red 40 Lake
Mfg. by Squire Boone Village
New Albany, IN 47150
For nutrition info, call 1-800-234-1804

Bases: Chemical Definition

Bases are anti-acids (antacids)

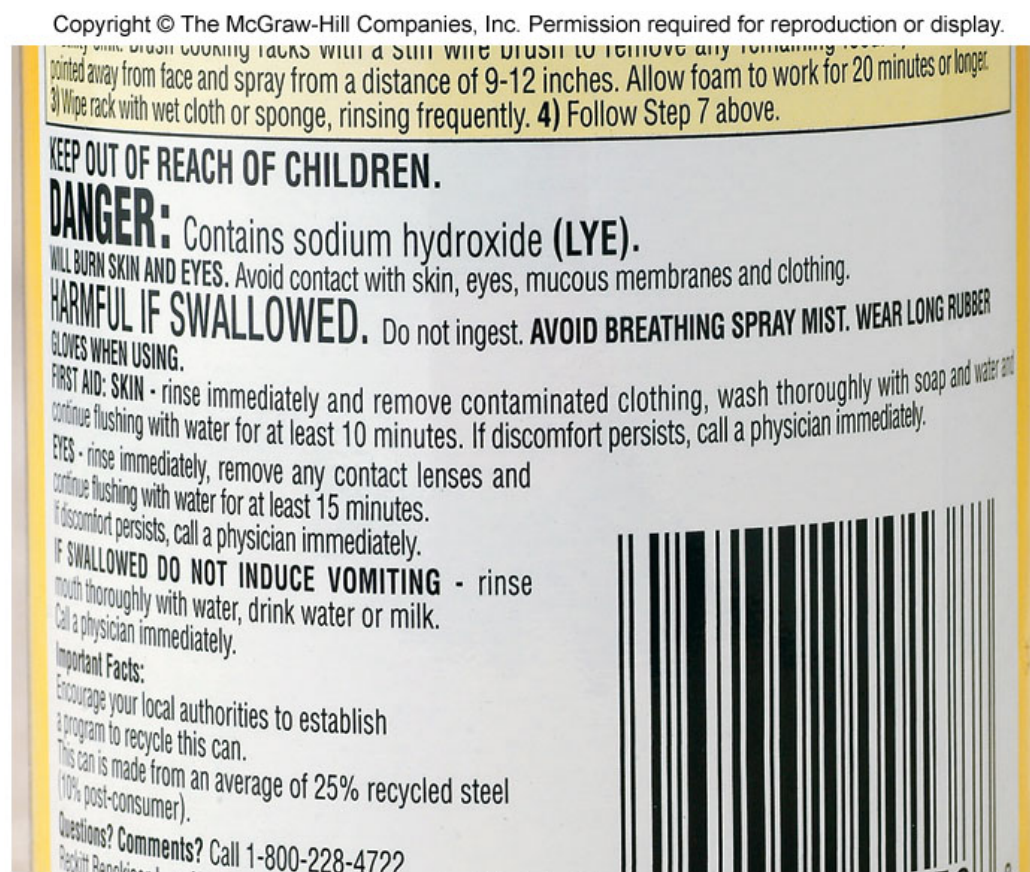
Bases produce **hydroxide anion** (OH^-) when dissolved in H_2O solution

In general: $\text{YOH} \rightarrow \text{Y}^+(\text{aq}) + \text{OH}^-(\text{aq})$

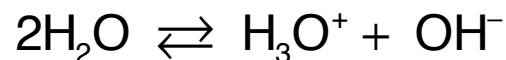
$\text{NaOH} \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$

$\text{NH}_3(\text{aq}) + \text{H}_2\text{O} \rightarrow \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq})$

Bases are bitter



Equilibria



- reaction proceeds in both directions at the same time
- adding more of a reagent will cause it to react faster
- this causes the equilibrium to shift towards the other side

$$K_{eq} = \frac{[\text{Products}]}{[\text{Reactants}]}$$

$$K_{\text{H}_2\text{O}} = \frac{[\text{H}_3\text{O}^+][\text{OH}^-]}{[\text{H}_2\text{O}]} \Rightarrow K_{\text{H}_2\text{O}} [\text{H}_2\text{O}] = [\text{H}_3\text{O}^+][\text{OH}^-] \Rightarrow 1 \times 10^{-14} = [\text{H}_3\text{O}^+][\text{OH}^-]$$

$$[\text{H}_2\text{O}] = \frac{\text{moles H}_2\text{O}}{\text{L solution}}$$

$$1 \text{ L H}_2\text{O} \times \frac{1000 \text{ g H}_2\text{O}}{1 \text{ L H}_2\text{O}} \times \frac{1 \text{ mole H}_2\text{O}}{18 \text{ g H}_2\text{O}} = 55.6 \text{ mole H}_2\text{O} \Rightarrow [\text{H}_2\text{O}] \text{ is } 55.6 \text{ M}$$

What is pH and Why Does it Matter?

pH stands for **power of hydrogen**

$\text{pH} = -\log[\text{H}_3\text{O}^+]$ (negative of the power of 10)

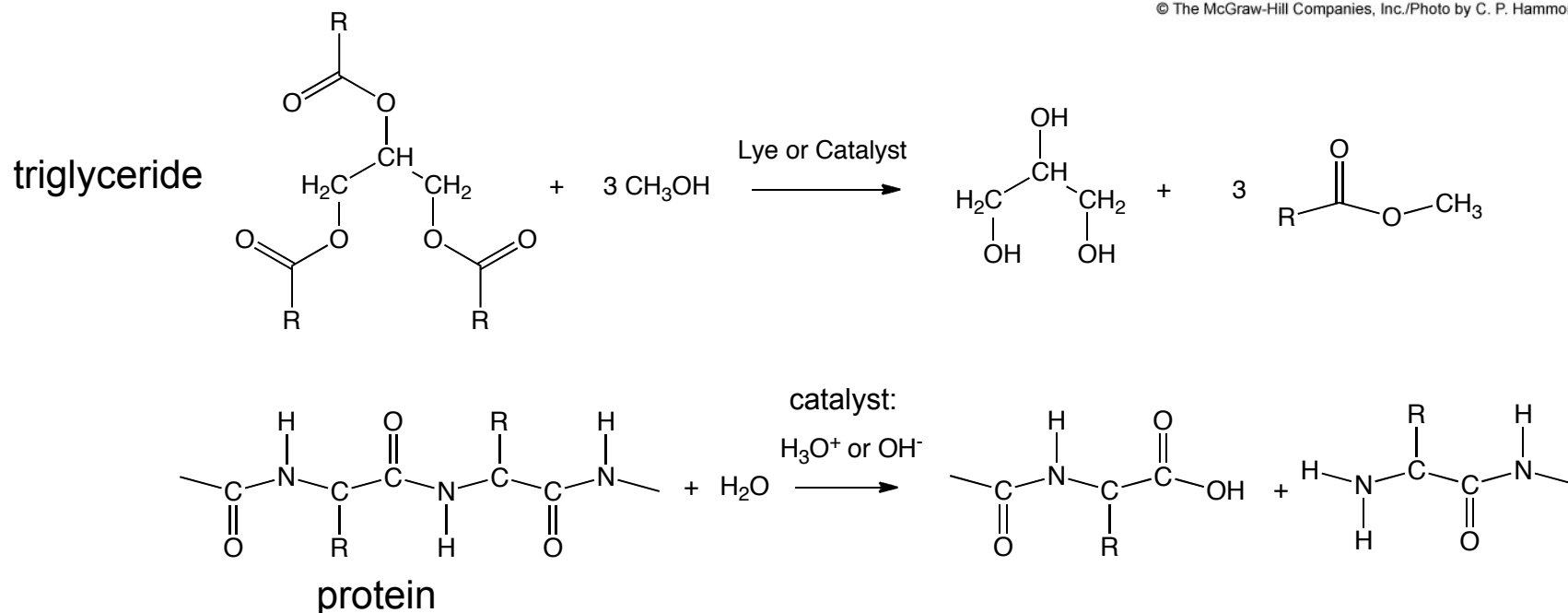
$[\text{H}_3\text{O}^+] = 10^{-\text{pH}}$

if $[\text{H}_3\text{O}^+] = 1 \times 10^{-4}$ then $\text{pH} = 4$

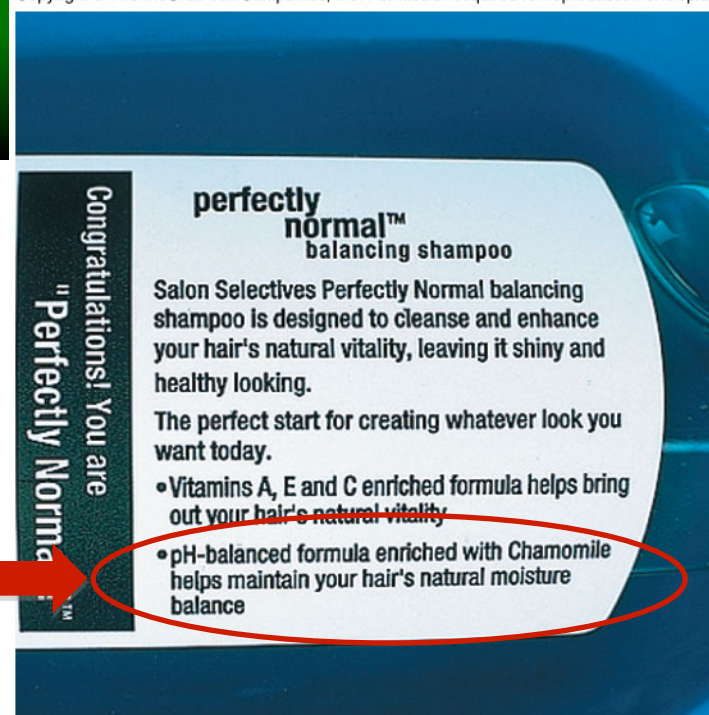
if $[\text{H}_3\text{O}^+] = 1 \times 10^{-9}$ then $\text{pH} = 9$

if $[\text{H}_3\text{O}^+] = 1 \times 10^{-7}$ then $\text{pH} = 7$

Acids & bases can break bonds in molecules of life



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